

Chemistry 1501, Dr. Hunter

Spring 2008

Exam # 2 (Group Part)

Name: \_\_\_\_\_, \_\_\_\_\_ Signature: \_\_\_\_\_

Name: \_\_\_\_\_, \_\_\_\_\_ Signature: \_\_\_\_\_

Name: \_\_\_\_\_, \_\_\_\_\_ Signature: \_\_\_\_\_

Name: \_\_\_\_\_, \_\_\_\_\_ Signature: \_\_\_\_\_

Name: \_\_\_\_\_, \_\_\_\_\_ Signature: \_\_\_\_\_

**Last name****First name**

The group portion of this exam has this title page plus three pages of questions. Please make sure you have all pages. Place the names (last name first) and signatures of each group member above. *Initial each page of the exam in the top right hand corner* using the initials of the all group members so that if your exam pages get separated I can match them to your group.

To obtain maximum credit for each question, show your work in detail. Partial credit for questions will not be assigned if no work is shown. **Indeed, no credit will be granted if complete work is not shown even for correct answers.** Feel free to use pictures/diagrams to illustrate your text answers and/or to use short text explanations to explain your drawings if your pictures are ambiguous. If you have to make assumptions, etc., to complete any answers, write me a short note stating and/or explaining your assumptions and testing them to the degree possible. Be very careful with significant figures on all answers. Note: Use the Atomic Masses from the table on the inside front cover of the text book that were also handed out with the exam as the front side of your Notes that you can bring to the exam (i.e., the back side of this 8.5 x 11" sheet).

You have 50 minutes for the group part of this exam. The 30 points for the individual part of this exam correspond to 15% of the 200 overall points for this course. Together, the group and individual parts of this exam are worth 50 points (i.e., 1/4 of the total course grade).

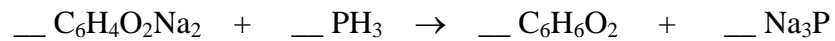
Grade

/30

(Group grade)

**G1.** (10 Points Total). Determine the number of moles of the indicated products and/or starting materials as requested:

(a) If one wants to produce 1.76 moles of  $\text{Na}_3\text{P}$ , how many moles of  $\text{C}_6\text{H}_4\text{O}_2\text{Na}_2$  should one use?



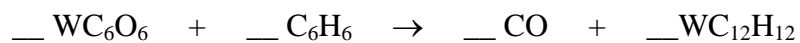
Grade	/10
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G2. (10 points total). Determine the number of grams of the indicated products and/or starting materials as requested:

(a) How many tonnes of Water are produced when one burns 32.16 kilograms of  $C_8H_{18}$ ?

Grade	/10
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G3. (10 points total). 12.2 moles of  $\text{WC}_6\text{O}_6$  were reacted with 28.4 moles of  $\text{C}_6\text{H}_6$ . Identify the limiting reagents and then predict the yields of Carbon Monoxide (i.e.,  $\text{CO}$ ) in moles.



Grade	/10
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